Self Check 4.3

Due No due date **Points** 7 **Questions** 7 **Time limit** None
Allowed attempts Unlimited

Instructions

This exercise will help you check your knowledge. Please take it as many times as you need to master the concepts. Select the best answer for each question.

Note: If there are short-answer or essay questions, your answer will automatically be marked correct regardless of how you respond, so please compare your answer with the feedback to make sure you know the correct answer!

Take the quiz again

Attempt history

	Attempt	Time	Score
KEPT	Attempt 2	15 minutes	5 out of 7 *
LATEST	Attempt 2	15 minutes	5 out of 7 *
	Attempt 1	20 minutes	0 out of 7 *

* Some questions not yet graded

ⓘ Correct answers are hidden.

Score for this attempt: **5** out of 7 *
Submitted 2 Apr 2019 at 10:14
This attempt took 15 minutes.

Question 1

1 / 1 pts

How much heat must be transferred by 100 g of liquid water at 0ºC in order to become ice? Is it absorbed or released?

- 80,000 cal
- 8,000 cal
- 100 cal
- 45,000 cal

Feedback: (100)(80) = 8,000 calories must be released

Question 2

1 / 1 pts

Is heat absorbed or released during condensation?

- released
- absorbed

Question 3

Not yet graded / 1 pts

When it snows outside, it usually feels warmer than right after the snow finishes. Explain this in terms of latent heat.

Your answer:

while it is snowing,latent heat is released .After it steps snowing ,there is no more latent heat being released.

The correct answer is "While it is snowing, latent heat is released (as the water freezes). After it stops snowing, there is no more latent heat being released."

Question 4

Not yet graded / 1 pts

When does the temperature of a substance stay the same even though heat is still being applied to it?

Your answer:

During a change of state.

The correct answer is "During a change of state."

Question 5

1 / 1 pts

Is heat absorbed or released during melting?

- (•) absorbed
- ( ) released

Question 6

1 / 1 pts

How much heat must be transferred for 40 g of liquid water at 100ºC to become steam?

- 20 cal
- 3200 cal
- 450 cal
- 21,600 cal

Feedback: (40)(540) = 21,600 calories must be absorbed

Question 7

1 / 1 pts

If 40 g of liquid at 100°C had heat transferred so it became steam, would you say heat was absorbed or released?

- released
- absorbed

Quiz score: **5** out of 7